## 大同大學 九十四 學年度 轉學考試 試題

考試科目:有機化學 系別:生物工程學系

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註:本次考試不可以參考自己的書籍及筆記; 不可以使用字典; 不可以使用計算器。

- 1. (a) Give the ground-state electron configuration for carbon (atomic number 6). (2%)
  - (b) How many electrons does carbon have in its valence shell? (2%)
  - (c) Determine the hybridization for the indicated atoms in each structure below. (6%)

- 2. Refer to the structure below to answer the following questions:
  - (a) Which of the labeled bonds in the structure are equatorial bonds? (2%)
  - (b) Which of the labeled bonds is *trans* to bond b? (2%)
  - (c) Which bonds have 1,3-diaxial interaction with each other? (2%)
  - (d) Draw the two chair conformations of cis-1-tert-butyl-4-chlorocyclohexane. Which is more stable? (4%)
- 3. Draw structures corresponding to the following names: (a) 3-Isopropylcyclopentene, (b) 1,4-Pentadiene, (c) 4-Chloro-2-heptyne, (d) p-Aminobenzoic acid, (e) 2,4,6-Trinitrotoluene (10%)
- 4. (a) Add curved arrows to the following reaction to indicate the flow of electrons: (5%)

(b) Follow the flow of electrons indicated by the curved arrows in the following reaction, predict the products that result: (5%)

5. Write the complete stepwise mechanism for the following reactions. Show all intermediate structures and all electron flow with curved arrows. (10%)

+ HBr -> (xcHzCH3 (b) (1) + CH3C-CL ALCL3 (-CH3 +

6. (a) Assign R or S configurations to the chirality centers in the following molecules: (4%)

- (b) Does cis-1,2-dimethylcyclobutane have any chirality centers? Is it chiral? Explain your answer. (6%)
- 7. Describe the effects of each of the following variables on S<sub>N</sub>2 reactions: (a) Solvent (b) Leaving group (c) Nucleophile (d) Substrate (10%)
- 8. Predict the major organic product(s) of the following reactions: (30%)

(a) CH3 1. Hg(0Ac)2, H20 (b) CH3 1. O3 (C)

2. NaBH4 (C) C=C-H 1. BH3, THF

2. H202, OH

(d) \( \frac{NB5}{CCL4} \) (e) \( \frac{OH\_1}{EHAN} \) (f) \(H\_3C - \frac{CH\_3}{CH\_3} \) \( \frac{Ne \ OCH\_2 OH\_2 OH\_3}{CH\_3 CH\_3 CH\_2 OH} \) (f) \( H\_3C - \frac{CH\_3}{CH\_3} \) \( \frac{Ne \ OCH\_2 OH\_2 OH\_3}{CH\_3 CH\_2 OH\_2 OH} \) \( H\_0\_2C - \frac{C}{C} \) CH\_3 \( \frac{F\_E Br\_2}{F\_E Br\_3} \)

(h) CH3 CCH2CH2COCH3 1. NaBH4

(i) CH3 CCH2CH2COCH3 2. H30 CH3 1. CH3 1. CH3 Mg Br

2. CH3 I. THF